**Explaining Polarity of molecules (Level 2) examiners tips: Read these please!**

• **1st**: state whether the molecule does or does not contain “polar bonds”

• if so, describe this in terms of the difference in “electronegativity” of the atoms in the molecule

*(please note you do not need to recall electronegativity numbers, don’t waste your time learning them!)*

• **2nd**: looking at the central atom state the number of “lone pairs” (if any) of the molecule

and (as a result of electronegativity and/or lone pairs) whether the molecule is “symmetrical” or not

• **3rd:** state the “shape” of the molecule

• **4th:** describe whether the dipoles or polar covalent bonds “cancel” or “do not cancel”and

• **finally**...an overall statement re polarity of the molecule...

*eg. “Molecules are polar if there is an uneven spread of charge over the molecule”*

*eg “Non-polar molecules have no dipole or polar bond present in the molecule or the spread of charge is*

*even”*

To achieve with an Excellence you must discuss fully the factors that affect both the shape and polarity of molecules

**Also…”don’t be daft”**

rush your thoughts about your answer, take your time and think before you write down your answer

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